**Describing type of solid, particle and bonding (Level 2) examiners tips: Read these please!**

• be logical here! look out for obvious ones first...

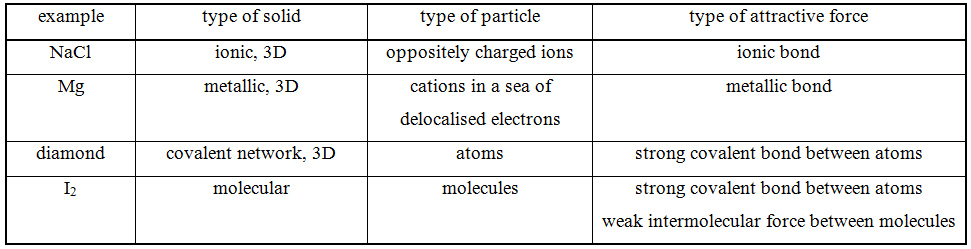
*eg. diamond, graphite and silica or silicon dioxide are covalent networks*

*eg metals are metallic bonding*

*eg metals bonded with non-metals are ionic*

*eg non-metals bonded to themselves are molecular substances*

• see the table below for a summary of the type of solid, particle and type of bonding/attractive force



**Also…”don’t be daft”, what NOT to do**

please note that there is no such thing as a molecular bond, no such thing!

Be aware that it is not good enough to state electrostatic attraction as the type of attractive force because all types of bonding are electrostatic attractions

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